

Thinking like a chemist 2/27/13

Calculating pH of a Buffer

- look at what you're given - draw a picture
- buffer acidic or basic:

Choosing a buffer

• Best situation:

- high concentration of conjugate acid/base partners
- one to one molar concentration will buffer against both acid and base
- one to one molar concentration

Acid/Base Titration

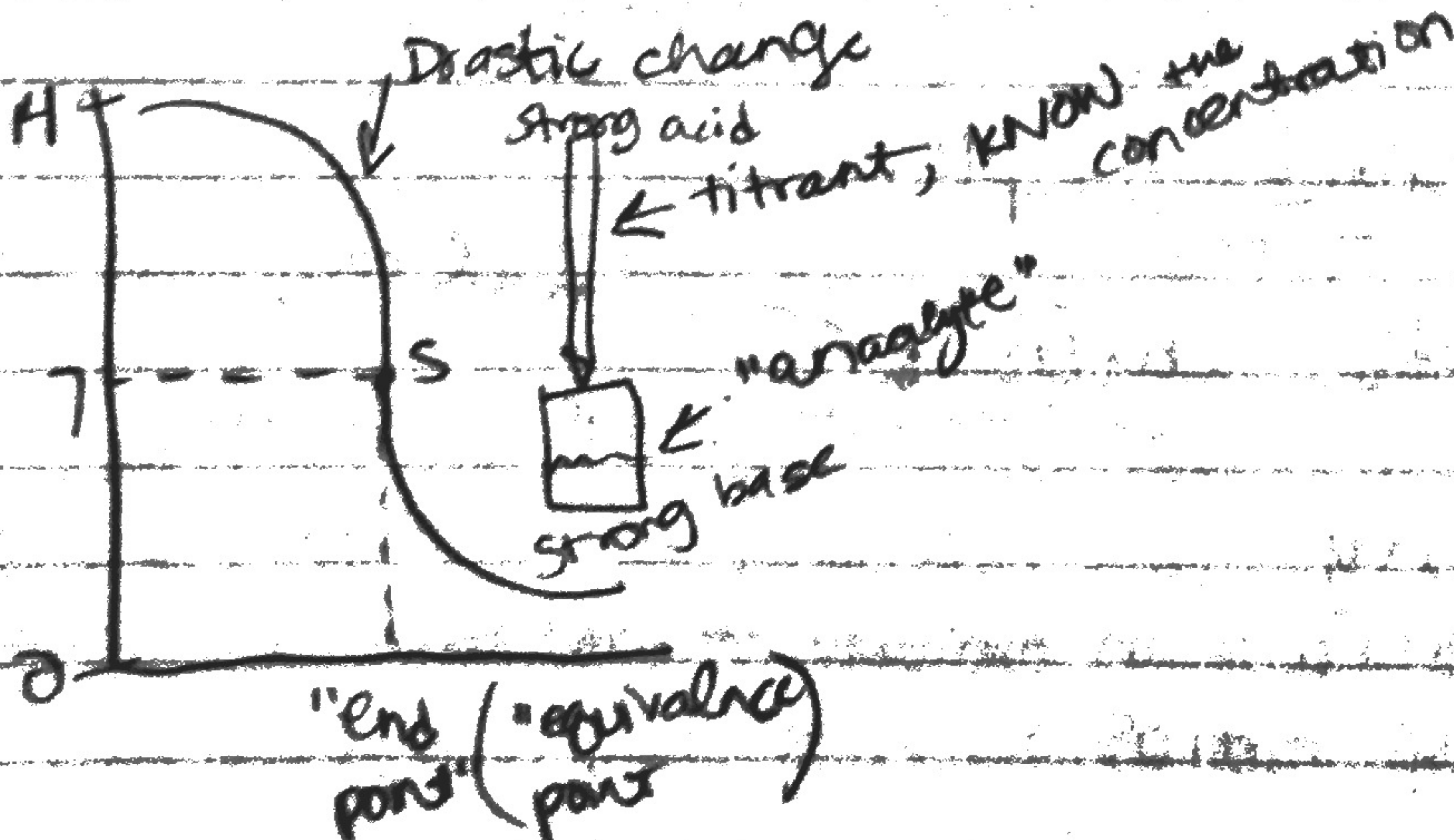
• Why do a titration?

- figure out the concentration

• You have a sol'n with an unknown property?

• unknown concentration?

• unknown K_a (K_b)?



• used to find endpoint

What do I care about the other protons?

- which = neutralize the acid

• As you take one proton off, the next one will come off

• If I add 0.1 moles of NaOH to 0.05 moles of H_3PO_4 , what will be the dominant species in solution?

• HPO_4^{2-}