UNIT6-DAY7-LaB1230pm

Wednesday, February 27, 2013
5:34 PM

Thinking Like a Chemist
About Acids and Bases
Part V

UNIT6 DAY7

CH302 Vanden Bout/LaBrake Spring 2013

What are we going to learn today?

Review Buffers
Explore Acid – Base Titrations
Explore the concept of protonation and pKa
Explore Behavior of Polyprotic Acids

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IMPORTANT INFORMATION

HW07 Due Tue 9AM
LM23 & 24 Due Tue 9AM

IMPORTANT INFORMATION

HW07 Due Tue 9AM
LM23 & 24 Due Tue 9AM

HW 6 → 100
- won't show up until END of unit

Exam 2 – Wed March, 6th - Rooms TBA
all new material done today

OPEN OFFICE HOURS FRIDAY 1-3PM GSB 2.126
(cookies)

Quiz: Clicker Question 1

Given a solution containing 1 M HF and 1 M NaF, and knowing that the pKa of HF = 3.14. The pH of the solution should be:

A) 10.86
B) 8.14
C) 7.00
D) 5.86
E) 3.14

Quiz: Clicker Question 2

Calculate the ratio of the molarities of acetate ions and acetic acid needed to buffer a solution at pH=5.25. The pK_a of CH_3COOH is 4.75.

$pH = pK_a + \log\frac{[A^-]}{[HA]}$

need more base

A) 1 to 3.2
B) .32 to 1
C) 3.2 to 1
D) 1 to .32

A) 1 to 3.2
B) .32 to 1
C) 3.2 to 1
~~D) 1 to .32~~

D) 3200 to 1

$pH = pK_a + \log\frac{[A^-]}{[HA]}$

need more base than acid

Calculating pH of a Buffer

Think First!

1. Look at what you are given, and think like a chemist.
2. Is the buffer acidic or basic?
3. Has additional acid or base been added to the buffer system?
4. If so, complete the neutralization reaction, and calculate the concentration of the buffer components
5. Choose the correct version of the Henderson-Hasselbalch equation.
6. Using the equation, calculate the pH.

All this is on website. Two worksheets for practice. Practice makes perfect.

Poll: Clicker Question 2

What is the purpose of a buffer?

One needs to keep an unusual microbial species alive in a laboratory setting. The microbe survives best in an alkaline environment with a pH > 9. The best choice of a buffering system would be equal molar amounts of:

A) $C_2H_5NH_2$, $C_2H_5NH_3^+$, $K_b = 5.6 \times 10^{-4}$
B) $C_6H_5NH_2$, $C_6H_5NH_3^+$, $K_b = 3.8 \times 10^{-10}$
C) $HClO_2$, ClO_2^-, $K_a = 1.2 \times 10^{-2}$
D) HOCl, OCl^-, $K_a = 3.5 \times 10^{-8}$

~4 $pK_b = 3.2$ $pK_a = 10$ pH ~ 10

$pK_b = 9.4$

$pK_a = 1.9$

A) $C_2H_5NH_2$, $C_2H_5NH_3^+$, $K_b = 5.6 \times 10^{-4}$
B) $C_6H_5NH_2$, $C_6H_5NH_3^+$, $K_b = 3.8 \times 10^{-10}$
C) $HClO_2$, ClO_2^-, $K_a = 1.2 \times 10^{-2}$
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$pK_b = 9.4$
$pK_a = 1.9$
$pK_a = 7.5$

Choosing a Buffer

The best situation is:

Relatively high concentration of conjugate acid-base partners.

One to one molar concentration will buffer against both added acid and added base.

One to one molar concentration – buffer will have a pH = pKa.

PRACTICE! HOMEWORK & WORKSHEETS!

$pOH = pK_b$

Acid Base Titration

Why do a titration?

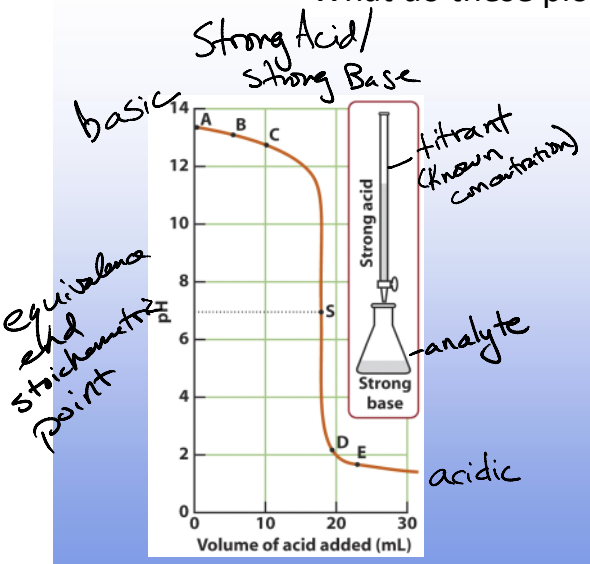
You have a solution with an unknown property

- Unknown Concentration?
- Unknown Ka (Kb)?

Both

Slowly neutralize the solution by adding a strong base (acid)
monitor the pH with each addition

What do these plots tell you?



Strong Acid/ Strong Base

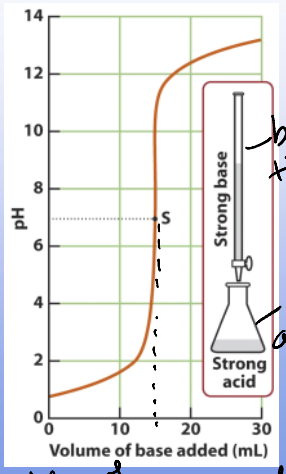
basic

titrant (known concentration)

equivalence and stoichiometric point

analyte

acidic

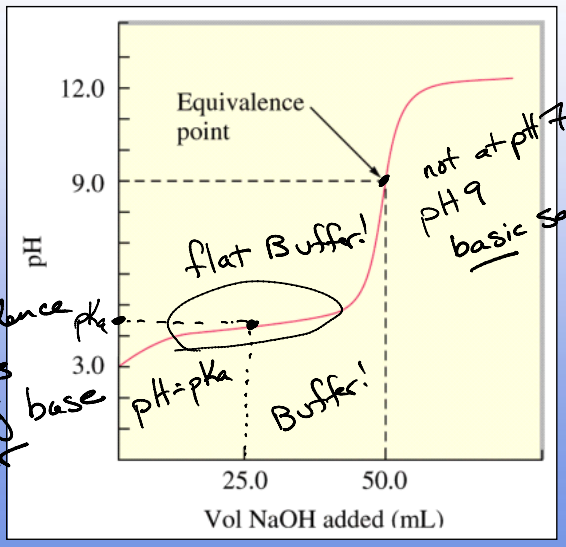


base titrant

acid analyte

Vol × mol/L = moles of acid

How is this plot different from previous plots?



NaOH

weak acid

not at pH 7

pH 9

basic salt

flat Buffer!

1/2 equivalence

equal amts acid & conj base

pH=pKa

Buffer!

partner

Work with neighbors on Titration Discovery Activity

Poll: Clicker Question 3

The initial concentration of the HBr is:

A) 7 M
B) .7 M
C) .007 M
D) .0007 M
E) 3.5 M

@ pH 7 have equal amts

0.7 mmol

0.7×10^{-3}

0.0007

0.12

Poll: Clicker Question 4

The pK_a for acetic acid is:

A) 1.7×10^{-5}
B) 1.6×10^{-9}
C) 4.76
D) 8.32
E) 3.43

Poll: Clicker Question 5

The initial concentration of acetic acid is:

A) 8 M
B) .8 M
C) .08 M
D) .008 M
E) .00008 M

What is the point of pH indicator?

http://youtu.be/Q21c3L3Ui8Q

http://youtu.be/IcU1YUGElLk

Demonstration:

Beaker containing _____ M HCl.

Add pH indicator.

Beaker containing _____ M NaOH.

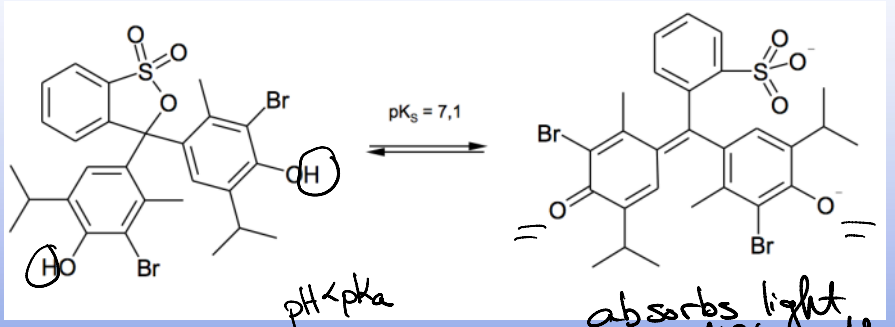
Observe.

Think of an explanation for why the indicator changes color.

Chemical Equilibrium
pH indicator

BTB indicator in acidic, neutral, and alkaline solutions (left to right).

Chemical Equilibrium
pH indicator Bromothymol Blue, pKa = 7.1



Protonated form
yellow
acid environment has protons

Deprotonated form
blue
basic solution does not have protons

Poll: Clicker Question 5

Chemical Equilibrium
pH indicator Bromothymol Blue

Bromophenol Blue has a pK_a of around 7. When it is protonated (HA form) it is yellow, when it is deprotonated (A^- form) it is blue.
What color would in be in a solution in which the pH was 9?

A. blue
B. yellow
C. green

basic

pH > pKa

deprotonated

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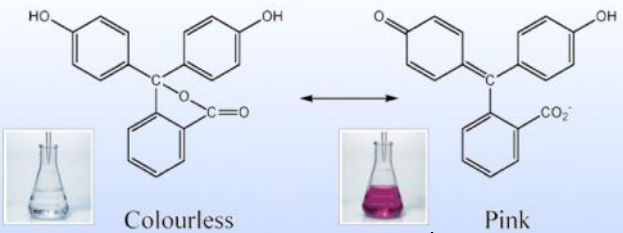
Indicator dye

Phenolphthalein

pKa = 8.2 $K_a = 6.3 \times 10^{-9}$



Colourless

Pink

low pH

h

$K_a = [H^+] \times \frac{[A^-]}{[HA]}$

$= [H^+] \times \frac{\text{Pink}}{\text{Clear}}$

HA

$[H^+] > 6.3 \times 10^{-9}$
pH < 8.2

$\frac{\text{Pink}}{\text{Clear}} < 1$

A^-

$[H^+] < 6.3 \times 10^{-9}$
pH>8.2

$\frac{\text{Pink}}{\text{Clear}} > 1$

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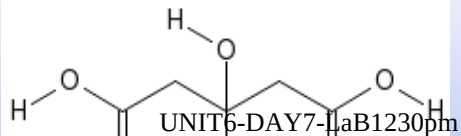
Polyprotic Acid

Each proton

3 Acid Protons



Polyprotic Acid

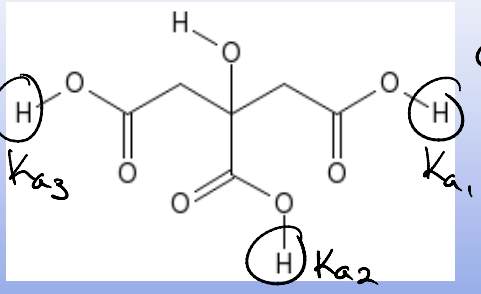
Each proton has its own Ka

1st H^+ comes off strongest

harder to pull off others

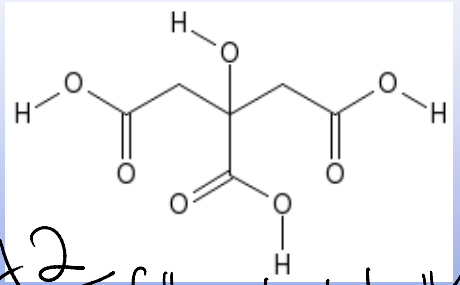
3 Acid

COOH



Poll: Clicker Question 7

Polyprotic Acid



pKa

$Ka_1 = 7.4 \times 10^{-4}$ 3.13

$Ka_2 = 1.7 \times 10^{-5}$ 4.77

pH of 5.5

$Ka_3 = 4.0 \times 10^{-7}$ 6.40

@pH 2 fully protonated pH < pKa

~~At pH = 2 will the acid be protonated or deprotonated?~~

~~a)Protonated~~

~~b)Deprotonated~~

c)~~Can't tell~~

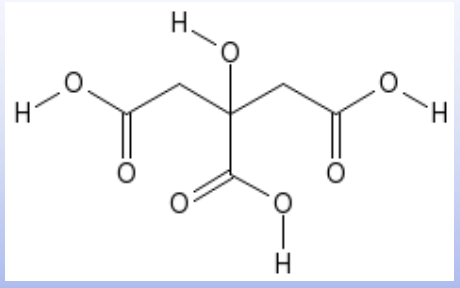
How many protons will be on the molecule?

A 1
B 2
C 3
D 0

② @ pH 5.5?

Poll: Clicker Question 8

Polyprotic Acid



$Ka_1 = 7.4 \times 10^{-4}$

$Ka_2 = 1.7 \times 10^{-5}$

$Ka_3 = 4.0 \times 10^{-7}$

At pH = 12 will the acid be protonated or deprotonated?
a)Protonated
b)Deprotonated
c)Can't tell

When do I care about the other protons?

When I neutralize the acid.

As you neutralize the first protons,
the second will come off,
....

If I add 0.1 moles of NaOH to 0.05 moles of H_3PO_4
what will be the dominant species in solution?

If I add 0.1 moles of NaOH to 0.05 moles of H_3PO what will be the dominant species in solution?

A. H_3PO_4 and $H_2PO_4^-$

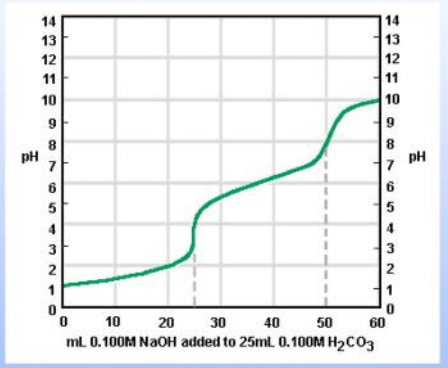
B. $H_2PO_4^-$

C. $H_2PO_4^-$ and HPO_4^{2-}

D. HPO_4^{2-}

E. HPO_4^{2-} and PO_4^{3--}

Titration of a polyprotic



Two equivalence points

Diprotic H_2A

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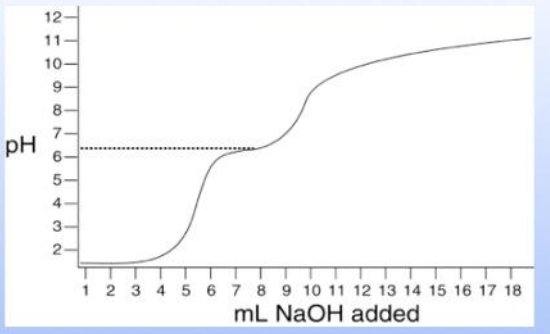
Given the following curve estimate K_{a2} for this unknown acid



at 1/2 equiv
pH = pKa
pH = 6.3
pKa = 6.3
$K_a = 5 \times 10^{-7}$

A. 1 B. 6.3 C. 5×10^{-6} D. 5×10^{-7}

Principles of Chemistry II

What did we learn today?

Acid Base titrations are performed to determine the concentration of an acid or base and/or the Ka of an acid or Kb of a base.

The pH can be calculated by determining the ending concentrations after a series of neutralization reactions are performed.

pKa indicates the extent of ionization of an acid
Lower the pKa value, the more acidic an environment necessary to keep the acid protonated